

APPLICATION FOR PATENT 1134.11A

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TITLE:

**ELECTROCHEMICAL CELL AND FUEL CELL WITH
CURABLE PERFLUORO-SULFONATE**

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SPECIFICATION

FIELD

5 **[0001]** The present invention relates to a curable liquid electrolyte for an electrochemical cell made of a perfluoro-sulfonate ionomer (PFSI), a cross linking agent, such as a vinyl monomer, and a protonic polymer.

BACKGROUND

[0002] The present application claims priority to co-pending U.S. Provisional Patent Application Serial No. 60/476,330 filed on June 6, 2003.

10 **[0003]** Polymer electrolyte membranes are useful in electrochemical devices such as batteries and fuel cells since they function as an electrolyte and also as a reactant separator. Typical membranes of this type are fabricated as thin films and then incorporated into cells and systems of various configurations.

15 **[0004]** Perflourinated hydrocarbon sulfonate ionomers Nafion™ by Dupont or analogous Dow perflourinated membranes are presently used as polymer electrolyte membranes for fuel cells. Such prior art membranes, however, possess severe limitations which, in addition to their cost, detract from their utility in new and advanced fuel cell designs.

20 **[0005]** Nafion™ is a costly material and difficult to replace due to its combined electrochemical performance and mechanical toughness. Other membranes possess compromising properties. For example, sulfonated polystyrenes degrade more rapidly than perflourinated hydrocarbon sulfonate ionomers (hereafter referred to as PFSI), sulfonated polyaromatics often adversely swell in water and form poor interfaces with catalyst layers as compared to PFSI. PBI/phosphoric acid
25 membranes do not operate below 120 degrees Celsius, and PFSI can operate below 120 degrees Celsius. It is hard to find a low cost alternative to PFSI because of the difficulties in finding suitable materials, synthetic methods, and the limited choice in

base materials.

5 [0006] However, while PFSI is so far the membrane material of choice, it is by no means ideal in its present condition. It is known that Nafion membranes soften only at temperatures >300 degrees Celsius. Additionally, sulfonated polyaromatics also possess very high softening temperatures and degrade while softening. Solid polymer electrolytes inherently possess very high softening temperatures by virtue of their high concentration of ions. Thus processing of membrane films to produce assemblies of membranes and catalyzed electrodes is restricted to high pressure, high temperature compression techniques which presently are a far from satisfactory situation.

10 [0007] A need has existed to find alternative materials capable of sustainable structures at high temperatures and pressures which are required to form electrode – electrolyte interfaces. As new applications for electrochemical devices emerge, and particularly as the scale of the devices is reduced, the limitations posed by the currently available electrolyte puts severe constraints on the feasibility of new designs.

15 [0008] A need has existed to reduce delamination of the membrane from the catalyst layer for electrode assemblies. Delaminating is due to poor chemical or physical bonding. It is hard to match the chemical nature of the membrane with the catalyst layer, or to form a chemical bond between the membrane and the catalyst layer. This failure mechanism is therefore a direct consequence of the manufacturing constraints imposed by the mechanical properties of the membrane material.

20 [0009] A need exists within the design of electrochemical cells for polymer electrolyte membranes having different properties than that of the PFSI materials. It would be desirable to have materials with similar electrochemical properties but which have different mechanical properties. Specifically, materials are needed that become soft and pliable at lower temperatures and that can be brought into intimate contact with catalyzed electrodes of arbitrary three dimensional shape without the use of high temperature compression methods.

5 [00010] The present invention is directed to a curable liquid electrolyte that provides options for assembly of electrochemical cells not currently available, allowing for fabrication of new topologies of electrochemical cells, and allowing the consideration of new materials for complementary components in electrochemical cells, that would potentially avoid major failure mechanisms of electrochemical cells and lead to overall reduced costs of both materials and manufacturing.

SUMMARY OF THE INVENTION

10 [00011] The invention relates to a fuel cell and electrochemical cell having a curable perfluoro-sulfonate ionomer based electrolyte composition. The composition is for nanosized cells with preferably between 10 wt% and 50 wt% of a perfluoro-sulfonate ionomer (PFSI) comprising acidic groups for transporting protons; between 10 wt% and 89 wt% of a monomer for dissolving the PFSI; and between 1 wt% and 60 wt% of a cross linking agent having at least two vinyl functionalities. Upon combining
15 the PFSI, monomer and cross linking agent, a curable electrolyte solution is formed with at least 50 wt% of the above components based on the total weight percent of the formed solution. Optionally, a quantity of initiator sufficient to cure the composition can be added to the composition, particularly if the electrolyte is to be cured using photo-curing or thermal curing.

20 [00012] The invention also relates to a method for producing a curable liquid electrolyte usable in nano-sized fuel cells and micro-structured electrochemical cells

BRIEF DESCRIPTION OF THE DRAWINGS

25 [00013] The present invention will be explained in greater detail with reference to the appended figures, in which:

[00014] Figure 1 schematically depicts an example of an electrochemical cell incorporating a membrane of the proton exchange membrane.

[00015] Figure 2 schematically depicts an example of a fuel cell of novel architecture requiring the use of non-conventional electrolytes.

[00016] Figure 3 depicts sulfonation level as a function of heating time for curable electrolyte mixtures based on sulfonated PEEK.

5 [00017] Figure 4 includes a graph indicating the relationship between time and sulfonation level of PEEK.

[00018] Figure 5 includes a table of proton conductivity values for cured liquid electrolytes.

[00019] Figure 6 includes a table of conductivity and water uptake values for liquid electrolytes with different sulfonation levels.

10 [00020] The present invention is detailed below with reference to the listed Figures.

DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENTS

[00021] Before explaining the present invention in detail, it is to be understood that the invention is not limited to the particular embodiments and that it can be practiced or
15 carried out in various ways.

[00022] The invention relates to fuel cells and other electrochemical cells using a curable perfluoro-sulfonate ionomer based electrolyte composition. The invention is particularly usable for nano-sized fuel cells that need very small particle sizes in the electrolyte. The composition has between 10 wt% and 50 wt% of a perfluoro-sulfonate ionomer (PFSI). The PFSI has acidic groups for transporting protons. The
20 composition also has between 10 wt% and 89 wt% of a monomer for dissolving the PFSI and between 1 wt% and 60 wt% of a cross linking agent having at least two vinyl functionalities.

[00023] Upon combining the PFSI, monomer, and cross linking agent, a curable electrolyte
25 solution is formed with at least 50 wt% of the above components based on the total

weight percent of the formed solution. The electrolyte is then usable between the oxidant and fuel plenums. The electrolyte is usable between the anode and the cathode of a fuel cell.

5 **[00024]** The composition can further include a quantity of initiator sufficient to cure the composition when using a procedure, such as photo-curing, thermal curing or such combinations. The initiator unexpectedly will facilitate curing without affecting catalyst used with the anodes or cathodes of the fuel cell. The faster curing time lowers the cost to make these types of nano-sized fuel cells.

10 **[00025]** A preferred composition contains between 20 wt% and 40 wt% of a PFSI comprising acidic groups for transporting protons, between 20 wt% and 70 wt% of a monomer for dissolving the PFSI, and between 5 wt% and 50 wt% of a cross linking agent having at least two vinyl functionalities. If an initiator is used, the composition includes between 5 wt% and 10 wt% of the initiator.

15 **[00026]** In an alternative embodiment, the electrolyte has between 10 wt% and 50 wt% of a perfluoro sulfonate ionomer (PFSI). The PFSI has acidic groups for transporting protons. The composition uses between 10 wt% and 89 wt% of a polar monomer and a polar solvent for dissolving the polar monomer. Between 1 wt% and 60 wt% of a cross linking agent having at least two vinyl functionalities is also used. The polar solvent is either water or an organic fluid. Some examples of usable polar
20 solvents include dimethylformamide, dimethylacetamide, n-methylpyrrolidinone, similar compounds, and combinations thereof.

[00027] When the PFSI, polar monomer, polar solvent, and cross linking agent are mixed into a formulation, a curable electrolyte solution is formed with at least 50 wt% of the above components based on the total weight percent of the formed solution.

25 **[00028]** The polar monomer is preferably a vinyl monomer bearing an acidic group, such as a sulfonic acid group, a phosphonic acid group, a carboxylic acid group, or combinations thereof.

5 **[00029]** The cross-linking agent preferably is a divinyl derivative of an organic compound. The organic compound can be an aliphatic, an aromatic, a heteroaromatic, or combinations thereof. The organic compound can also be sulfonic acid, a sulfone, a phosphate, a phosphone, a phosphonic acid, a carboxylate, a carboxylic acid, an acrylate, a methylacrylate, an acrylamide, a methacrylamides, or combinations thereof.

10 **[00030]** The cross linking agent preferably is a trivinyl derivative of an organic compound. The organic compound for the cross-linking agent can be sulfonic acid, a sulfone, a phosphate, a phosphone, a phosphonic acid, a carboxylate, a carboxylic acid, an acrylate, a methylacrylate, an acrylamide, a methacrylamides, or combinations thereof.

15 **[00031]** The invention also contemplates that the curable liquid electrolyte solution can further comprise an elastising agent. The elasticizing agent can be a polymerizable vinyl monomer to enhance the toughness of structure of the cured electrolyte.

20 **[00032]** Micro and nano structured materials refer to materials which have feature sizes that are measured in microns or smaller. The feature sizes can be a benefit of the material structure itself, as is the case of materials which are porous with very small pore sizes. Alternatively the features can be formed into the material to create a patterned microstructure.

25 **[00033]** Micro-structured materials have many practical applications including filtration, production of electronic components, micro-sensors, fuel processing, and electrochemical cells.

30 **[00034]** Non-porous materials, into which micro-structures are commonly formed, can be a silicon wafer, an alumina, various ceramics, metals and plastics. These materials are processed using methods, such as laser ablation, wet etching, deep reactive ion etching, electro-discharge machining, dicing, water jet cutting, micro-injection molding, casting or molding.

35 **[00035]** Porous media with defined micro-structures typically have pore sizes ranging from

less than 1nm to 10 microns, more preferably between 20 nm and 100 nm. Porous media with micro-pores can be an aero-gel, a foam, or a mat or felt of conductive and non-conductive materials. The porous media can be porous plastics and ceramics.

5 **[00036]** In electrochemical cell applications, including fuel cells, it is desirable to combine the micro-structured porous material properties with the micro-structural formation to integrate a three-dimensional topology with the natural micro-structure of the porous material. Such integrated structures can be formed from monoliths of porous media by machining, by molding, or by compressing or sintering micro-powders into
10 a mold to create the desired micro-structured, micro-porous shape.

[00037] Nano-structured materials are porous materials with features sizes measured on the order of nano-meters or Angstroms. Nano-structured materials are formed in a similar manner to the micro-structured materials.

[00038] Creation of complex microstructures using micro-structured materials presents
15 challenges for the creation of composite structures. Composite structures can be formed by the deposition of one or more metals onto or into a microstructure to render the surfaces of the microstructure catalytically active. The deposition of conductive materials onto a microstructure renders the structure conductive. Graded porosity is created by selectively filling micro-pores with a barrier material or the
20 deposition of micro-particles of micro-porous materials onto some macro-porous carrier.

[00039] In all these applications, the deposition of nano-particles and nano-crystals onto a microstructure is difficult using conventional particle deposition techniques. Electro-phoretic deposition and electroplating cannot work with nano-sized particles,
25 physical vapour deposition techniques can coat the visible surfaces only, with a uniform film, and do not provide the opportunity for mixing multiple materials on the micro-structured surface. Furthermore, vapour deposition methods cannot penetrate deeply enough into a structures micro-features.

[00040] The method of depositing nano-particles and nano-crystals directly onto and into microstructures uses high concentration nano-particle dispersions. These dispersions have been formulated to provide very stable high concentration dispersions of nano-particles or nano-crystals in a carrier liquid having very low overall viscosity. The colloidal dispersion is then easily deposited on a microstructure whereupon the capillary interaction of the microstructure and the colloidal dispersion pulls the liquid into the structure. The suspended nano-particles then interact with the microstructure and are pulled out of the dispersion. Heat treating or otherwise post-treating the deposited particles strengthens the adhesion of the particles to the substrate.

[00041] The very small feature sizes of the micro-structured material, whether they are porous or non-porous makes them behave like capillaries. Before providing an example of the method, we will explain the theory of capillary action.

[00042] The spontaneous flow of a liquid through a capillary is normally described in terms of the Washburn equation:

$$\frac{h^2}{t} = \frac{r\gamma_{LV} \cos\Theta}{2\eta}$$

where t is the time needed for a liquid to reach the penetration height/depth of h , r is the capillary radius, γ_{LV} is the liquid surface tension, η is the liquid viscosity, and Θ is the three-phase contact angle between the liquid, its saturated vapor and capillary wall.

[00043] All porous materials are conveniently treated as consisting of bundles of capillaries that can be characterized by some effective radius r_{eff} given by the following equation:

$$r_{eff} = \frac{2(1-\phi)}{\phi\rho_s A}$$

where ϕ is the volume fraction of solid in the porous material, ρ_s is the density of solid, and A is the specific surface area per gram of solid.

[00044] As the Washburn equation indicates, the penetration rate reaches a maximum value when the liquid completely wets the capillary walls ($\Theta = 0$). In the case of hydrophobic solids, such as porous carbons, characterized by large water contact angles ($\Theta \approx 90$ deg), penetration rates are extremely low but this effect can partly be compensated by using a material with a higher value of r_{eff} , i.e. a material with a small internal specific surface area A . Such capillary size effects were clearly observed in studies using materials with varying porosity. While the deposition of platinum nano-particles from aqueous dispersions was almost impossible in highly porous and hydrophobic substrates, the same type of substrate with a lower porosity could easily be saturated with aqueous platinum nano-dispersions.

[00045] Highly hydrophobic solids are easily wetted ($\Theta = 0$) by organic solvents with low surface tensions γ_{LV} , such as the solvent methanol. Therefore, the penetration rates of highly porous and strongly hydrophobic substrates can significantly be increased using water/methanol mixtures. Platinum catalyst particles are dispersed in a polymer solution. The polymer is absorbed onto the internal surfaces of the hydrophobic substrate during deposition rendering these surfaces strongly hydrophilic ($\Theta = 0$), thus dramatically increasing the penetration rates. This allows the catalyst to be deposited uniformly within a certain volume of the substrate.

[00046] Example. In fuel cells, an electrochemical cell is formed by the association of an electrolyte with catalyzed electrodes. Typically, these electrodes are formed by the deposition of supported catalysts onto a porous gas diffusion electrode. Typical pore sizes are larger than 2 microns, with catalyst supported on the micro-porous materials in particles several microns in diameter.

[00047] Fuel cells can support the use of micro-structured gas diffusion media directly, but the use of the supported catalyst is problematic in such cases as the size of the supported catalyst particles is too large compared to the feature size of the microstructure.

5 [00048] The invention produces an electrochemical cell with a curable liquid electrolyte. The method begins by mixing a perfluoro sulfonate ionomer (PFSI) solution having a solvent with a monomer and a cross linking agent having at least two vinyl functionalities forming a mixture. The solvent of the PFSI solution is removed by distillation to obtain a curable PFSI electrolyte. The curable PFSI electrolyte is disposed on a substrate forming an intermediate.

10 [00049] The method then ends by treating the intermediate to cure the curable PFSI electrolyte into a cured electrolyte with increased viscosity and increased structural strength. The curable PFSI electrolyte can also be treated into a gel, a solid, a liquid with a higher viscosity than the mixture, or combinations thereof.

[00050] The invention contemplates that the method further includes the step of adding an initiator to the mixture and treating the intermediate with a procedure, such as photo-curing, thermal curing and combinations thereof.

15 [00051] The preferred method for treating the intermediate is by electron bombardment. The electrochemical cell is then constructed by disposing a catalyst, such as a noble metal, a transition metal, an alloy of noble metals, an alloy of transition metals, or combinations therein, on a substrate.

20 [00052] The substrate in the method is a mold for forming a shape, a porous media, an electrode, a catalyst support, laminates thereof, and composite structures of plastics, graphics, ceramics, glass, silicon, structural metals or combinations thereof. If porous media is used, the porous media is a foam, an aero gel, a mat, felt, paper, mesh, laminates thereof, composites thereof or combinations thereof.

25 [00053] The electrochemical cell of fuel cell can be formed by disposing a catalyst on the substrate prior to disposing the curable PFSI electrolyte on the substrate. At least a portion of the curable PFSI electrolyte may be removably disposed on the substrate.

[00054] The method can also include the step of adding a solvent to the mixture or adding an elastising agent to the mixture for the electrolyte. An example of an elasticizing agent

is a polymerizable vinyl monomer to enhance the toughness of structure of the cured PFSI electrolyte.

5 [00055] Now and with reference to the Figures, FIG 1 schematically shows an example of an electrochemical cell incorporating a membrane of the proton exchange membrane of the invention.

10 [00056] Figure 1 shows a cross section of a conventional electrochemical cell in which a solid polymer electrolyte sheet (10) is bonded to a pair of catalyzed gas diffusion electrodes (20 and 30). The assembly that results is known as a membrane electrode assembly (MEA) (40). The MEA (40) is then clamped between flow plates (50 and 60) to form a working electrochemical cell (70), in this case a fuel cell. In this example, pre-formed electrolyte sheets (10) interface between the electrolyte (10) and the electrodes (20 and 30) and rely on the mechanical bonding of the discrete components. Variations in the approach to manufacturing cells this way are used, but most continue to rely upon the mechanical bonding of the discrete MEA components. The present invention does not uses ionic bonding, rather than mechanical bonding,

20 [00057] The electrolyte sheet (10) can be formed using the present invention by casting the electrolyte and assembling a cell. In a preferred embodiment, electrolyte sheet (10) can be formed directly on either electrode (20 or 30), thereby forming a more intimate contact between the electrolyte and one of the electrode surfaces, which is advantageous in reducing contact resistance and improving the catalyst-electrolyte-conductive media interface.

25 [00058] In another preferred embodiment, the electrolyte sheet (10) can be formed by applying curable liquid electrolyte according to the present invention between two electrodes (20 and 30) and then curing the resulting assembly. Curing can be by thermal, photo or electron beam methods depending upon the cross-linking initiator used in the mixture. By assembling the membrane electrode assembly in this manner, the improved electrode-catalyst interface is obtained for both electrolyte – electrode interfaces. In addition, the formation of the membrane electrode assembly

in this manner does not require high temperatures or high pressures to be employed, allowing for flexibility in the selection of materials for porous electrodes (20 and 30) and lower manufacturing costs

5 [00059] Figure 2 shows a cross sectional view of a micro fuel cell (100) in which curable electrolyte layers are applied by spin-coating the electrolyte (110) onto a planar substrate (120) containing multiple anode regions (130). As with the previous example, the uncured electrolyte can be used as the basis for assembly of cathodes (140) onto the individual fuel cells. Due to the properties of the curable liquid electrolyte the dimensions of the spin-coated electrolyte layer will not alter significantly during curing.

10 [00060] Figure 3 shows the cross-sectional view of yet a different fuel cell topology in which a fuel cell (200) is created by forming a channel (210) within a porous substrate (220). Catalyst (230) is added to create anode (240) and cathode (250) regions. Finally, the electrolyte (260) is injected into the channel and then cured to form the electrolyte layer in-situ within the fuel cell. It would be difficult to assemble this fuel cell without a form of curable liquid electrolyte. Solvent based liquid polymer electrolyte formulations contain very low percentage of electrolyte and would shrink significantly as the solvent is driven off in order for the electrolyte to solidify.

15 [00061] This example discloses the preparation of liquid electrolyte using sulfonated PEEK as an example of an ionic polymer. Sulfonated PEEK was prepared by direct sulfonation reaction of PEEK with concentrated sulfuric acid (95~98%). An amount of PEEK (18 grams) was first dissolved in 300 ml sulfuric acid at room temperature. The reaction was carried out overnight at room temperature, followed by heating for 3 hours to 8 hours on a 55 degrees Celsius oil bath to control the sulfonation level of resultant polymers.

20 [00062] Figure 4 shows the time dependence of sulfonation levels at 55 degrees Celsius. The ability to control sulfonation is important for control of eventual electrolyte performance characteristics.

5 [00063] Liquid electrolyte was prepared by dissolution of S-PEEK in a vinyl monomer mixture in the absence or presence of a photo-initiator or a photo-initiating system. Generally, vinyl monomers mixtures consist of a protonic monomer, vinyl phosphoric acid, cross-linking agent divinyl sulfone, and a third monomer acrylonitrile. In some cases, water and or an organic compound, N, N-dimethylacetamide, (DMA) can be added. The liquid electrolyte was cured by photo or electro beam (EB) exposure.

10 [00064] A second example of a photo-curable liquid electrolyte uses 30% wt/wt S-PEEK (71% sulfonation) and 70% monomer mixture (22.5% vinyl phosphoric acid, 52.5% divinyl sulfone, and 25% acrylonitrile of the mixture) in the presence of 3.75% photo-initiator of monomers was formed with good mechanical properties and good proton conductivity usable for a fuel cell or other electrochemical cell..

15 [00065] This electrolyte used S-PEEK (865 mg) dissolved in 2000 mg monomer mixture composing of vinyl phosphoric acid (450 mg), divinyl sulfone (1050 mg), and acrylonitrile (500 mg) in the presence of a photo initiator (75 mg). An electrolyte membrane was prepared by casting the resultant liquid electrolyte on a glass slide and exposing the film to visible light for four hours. The membrane was freed from the slide by soaking with water. Proton conductivity was 0.058 S/cm at room temperature and 100% humidity. The conductivity measurement was made with an impedance analyzer.

20 [00066] An electron beam curable liquid electrolyte composed of 30% wt/wt S-PEEK (58% sulfonation) and 70% monomer mixture (15% vinyl phosphoric acid, 55% divinyl sulfone, and 30% methylacrylonitrile of the mixture) in the presence of 12.2% of N, N-dimethylacetamide was made.

25 [00067] Another example of a curable electrolyte is S-PEEK (865 mg) dissolved in 2000 mg monomer mixture comprised of vinyl phosphoric acid (300 mg), divinyl sulfone (1100mg), and methylacrylonitrile (600 mg). An electrolyte membrane from this mixture was prepared by casting the resultant liquid electrolyte on a glass slide forming a film and exposing the electrolyte to a electron beam with a dose of about

250 KGy. The proton conductivity of the membrane is 0.060 S/cm at room temperature with 100% humidity.

5 [00068] In a preferred cell, small voids can be formed directly within the liquid electrolyte that acts as a water reservoir to aid in humidification of the cured membrane. These voids are formed by including a high boiling point organic compound within the liquid electrolyte mixture. A liquid electrolyte composing 30% wt/wt S-PEEK (63% sulfonation) and 70% monomer mixture (15 vinyl phosphoric acid, 45% divinyl sulfone, and 30% acrylonitrile of the mixture) in the presence of 20% of N, N-dimethylacetamide was usable.

10 [00069] In still another example, S-PEEK (865 mg) was dissolved in 2000 mg monomer mixture composing of vinyl phosphoric acid (300 mg), divinyl sulfone (900 mg), and acrylonitrile (800 mg) in the presence of a photoinitiator (75 mg) and N, N-dimethylacetamide (567 mg). An electrolyte membrane was prepared by casting the resulting liquid electrolyte on a glass slide and exposing the film to visible light for 4
15 hours. Nano-voids were created by removal of N, N-dimethylacetamide under vacuum at room temperature.

[00070] In another version of this example, the temperature could be elevated to drive off the N, N-dimethylacetamide. The proton conductivity of the resultant membrane was found to be 0.039 S/cm at room temperature with 100% humidity.

20 [00071] In still another embodiment, water is included within the liquid mixture described above so that after curing a pre-hydrated membrane is formed. The pre-hydrated membrane preferable made from a liquid electrolyte composed of S-PEEK (71% sulfonation) and a monomer mixture (22.5% vinyl phosphoric acid, 52.5% divinyl sulfone, and 25% acrylonitrile of the mixture) in the presence of water. For this
25 example, S-PEEK (865 mg) was dissolved in 2000 mg monomer mixture composed of vinyl phosphoric acid (450 mg), divinyl sulfone (1050 mg), and acrylonitrile (500 mg) in the presence of a photoinitiator (75 mg) and water (600 mg). An electrolyte membrane was prepared by casting the resultant liquid electrolyte on a glass slide forming a film and exposing the film to visible light for three hours. The proton

conductivity of the membrane was measured at 0.060 S/cm at room temperature with 100% humidity.

[00072] Figure 5 shows a table of performance data indicating that the conductivity of the cured electrolyte can be controlled by controlling the sulfonation level of the SPEEK as previously described. The figure provides measured results for 3 liquid electrolytes composing 30% wt S-PEEK with different sulfonation levels and 70% monomer mixture (15 vinyl phosphoric acid, 42.5% divinyl sulfone, and 42.5% acrylonitrile) in the presence of 2.6% photoinitiator and 20% of N, N-dimethylacetamide.

[00073] For these examples, S-PEEK (865 mg) was dissolved in 2000 mg monomer mixture composed of vinyl phosphoric acid (300 mg), divinyl sulfone (850 mg), and acrylonitrile (850 mg) in the presence of a photo initiator (75 mg) and N, N-dimethylacetamide (567 mg). Electrolyte membranes were prepared by casting the resultant liquid electrolytes on glass slides forming a film and exposing the films to visible light for four hours. The proton conductivity of those membranes at room temperature with 100% humidity is shown in Figure 5.

[00074] The conductivity and the amount of water uptake of the electrolyte can be controlled by controlling the degree of cross-linking when the curable liquid electrolyte is cured. A liquid electrolyte which has the composition of 30% S-PEEK (63% sulfonation) and 70% monomer mixture (30:62.5:22.5 ratio of vinyl phosphoric acid, divinyl sulfone, and acrylonitrile) can be usable. This electrolyte forms S-PEEK (865 mg) dissolved in a monomer mixture (200 mg) composing of vinyl phosphoric acid (300 mg), divinyl sulfone (850 mg), and acrylonitrile (850 mg) in the presence of a photo initiator (75 mg) and N, N-dimethylacetamide (567 mg). Electrolyte membranes were prepared by casting the resultant liquid electrolyte on glass slides forming a film and exposing them to visible light for four hours. The resultant membrane showed high mechanical strength and high stability in water due to the high degree of cross-linking resulting from the large amount of divinyl sulfone used in the mixture.

[00075] Figure 6 demonstrates the room temperature proton conductivity and water uptake of this membrane and another membrane prepared from a similar liquid electrolyte using a lower content of divinyl sulfone.

[00076] Conductivity and water uptake are also controllable by controlling the degree of polymerization. A liquid electrolyte composing of 70% S-PEEK (58% sulfonation) and 70% monomer mixture (30:62.5:22.5 ratio of vinyl phosphoric acid, divinyl sulfone, and acrylonitrile) with 14% of N, N-dimethylacetamide was cured by electron beam with different doses to control the degree of polymerization of monomers.

[00077] For this version. S-PEEK (865 mg) was dissolved in a monomer mixture (200 mg) composing of vinyl phosphoric acid (300 mg), divinyl sulfone (1300 g), and acrylonitrile (400 mg) in the presence of N, N-dimethylacetamide (400 mg). Electrolyte membranes were prepared by casting the resultant liquid electrolyte on a glass slide forming a film and exposing to an electron beam with doses varying from 150 KGy to 350 KGy.

[00078] The examples of curable liquid electrolytes given so far have been based on S-PEEK. In another embodiment, curable liquid electrolytes are formed based on the use of recast NafionTM for nano-sized fuel cells.

[00079] A curable recast NafionTM electrolyte was prepared by exchanging the original solvents in 5% NafionTM 117 solution with vinyl monomers. NafionTM solution was added to a vinyl monomer mixture in the absence or presence of a photo initiator or a photo initiating system. The vinyl monomer mixture consisted of a protonic monomer, a vinyl phosphoric acid, a cross-linking agent, and a divinyl sulfone. A high boiling point organic solvent, N, N-dimethylacetamide, (DMA) may be added for enhanced annealing of the cured electrolytes. The original solvents of the NafionTM solution were removed by soft vacuum distillation using a rotary evaporator to give a photo-curable or electron-beam-curable liquid electrolyte. Cured electrolytes were annealed between 140 degrees Celsius and 160 degrees Celsius to improve their mechanical strength and solvent resistance.

[00080] In another embodiment, a photo-curable recast NafionTM electrolyte composed of 30% Nafion 117 and 70% monomer mixture (15:65:20 ratio of vinyl phosphoric acid, divinyl sulfone, and acrylonitril) was prepared. The 5% NafionTM 117 solution was mixed with a monomer mixture of vinyl phosphoric acid (300 mg) and divinyl sulfone (1300 mg) in the presence of a photoinitiator (75 mg) and N, N-dimethylacetamide (500 mg). After the removal of the original solvent, acrylonitrile (400 mg) was added. A membrane was prepared by casting a liquid layer on a glass slide forming a film and curing with a visible light for four hours. The membrane was annealed on the glass slide between 140 degrees Celsius and 150 degrees Celsius for 4 hours in a nitrogen atmosphere. A strong and flexible recast NafionTM membrane was freed from the glass by soaking with deionized water. This membrane shows excellent stability in boiling water with a water uptake of 31% at room temperature.

[00081] In an alternate embodiment, an electron-beam curable recast NafionTM electrolyte was formed composed of 30% NafionTM 117 and 70% monomer mixture (15:55:30 ratio of vinyl phosphoric acid, divinyl sulfone, and methylacrylonitrile) in the presence of N, N-dimethylacetamide. The 5% NafionTM 117 solution was mixed with a monomer mixture of vinyl phosphoric acid (300 mg) and divinyl sulfone (1100 mg) in the presence of N, N-dimethylacetamide (500 mg). After the removal of the original solvents, methylacrylonitrile (600 mg) was added. A membrane was prepared by casting a liquid layer of the electrolyte on a glass slide and curing with electron beam (EB) at doses of about 100 KGy. The membrane was annealed on the glass slide between 140 degrees Celsius and 150 degrees Celsius for four hours in a nitrogen atmosphere. The proton conductivity of the cured membrane is 0.031 at room temperature.

[00082] The conductivity of the recast Nation liquid electrolytes can be controlled by controlling the degree of cross-linking. A curable recast NafionTM electron composed of 30% NafionTM 117 and 70% monomer mixture (15:42.5:42.5 ratio of vinyl phosphoric acid, divinyl sulfone, and acrylonitrile) was prepared with a lower degree of cross-linking than the original recast NafionTM example. NafionTM 117

solution was mixed with a monomer mixture of vinyl phosphoric acid (300 mg) and divinyl sulfone (850 mg) in the presence of a photo initiator (75 mg) and N, N-dimethylacetamide (500 mg). After the removal of the original solvents, a third monomer, acrylonitrile (850 mg) was added. A membrane was prepared by casting a liquid layer on a glass slide and curing with visible light for four hours. The membrane was annealed on the glass slide between 140 degrees Celsius and 150 degrees Celsius for four hours in a nitrogen atmosphere.

[00083] While this invention has been described with emphasis on the preferred embodiments, it should be understood that within the scope of the appended claims, the invention might be practiced other than as specifically described herein.